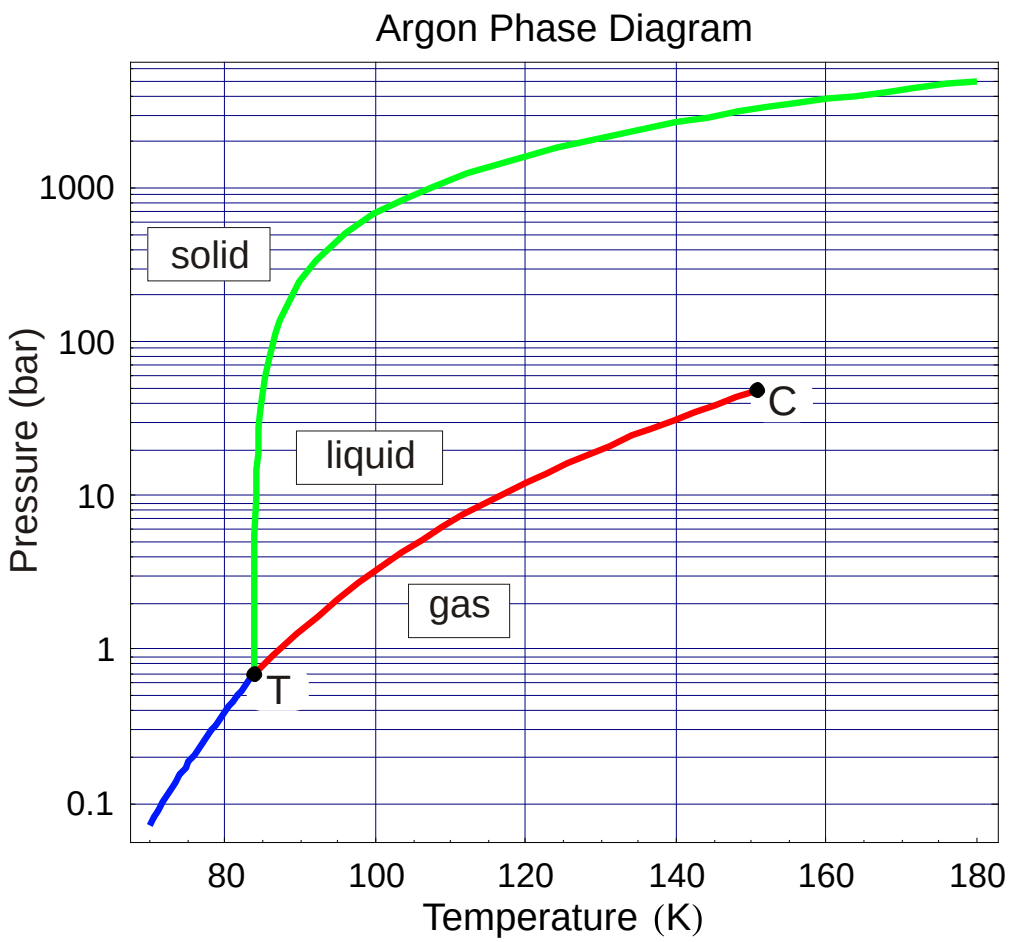
Thermodynamic Properties

Phase Diagram [4]



Enthalpy

Fit to enthalpy data from [5]

Pade approximant:

$$\Delta H = \frac{A + BT}{1 + CT}$$

ΔH = enthalpy of vaporization in kJ/mol

T = temperature in K

for $83.8 \le T \le 100$, $A = 7.98304$, $B = -0.0481275$, $C = -0.0047259$

RMS deviation = 0.167 J/mol